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| **Keyword** | **Definition** |
| **Molecule** | A group of atoms joined together. Most non-metals consist of molecules. Most compounds of non-metals also consist of molecules. |
| **Atmosphere** | The layer of gases that surround the Earth. |
| **Mixture** | Two or more different chemicals, mixed but not chemically joined. |
| **Particles** | Tiny bits of a substance. |
| **Condensed** | The change of state of a substance from a gas to a liquid.  |
| **Photosynthesis** | A chemical reaction in green plants that uses the energy in sunlight to make glucose and oxygen. |
| **Sedimentary rock** | Rocks formed in layers over time. |
| **Pollutant** | Waste matter that contaminates soil, water or air. |
| **Emissions** | Something given out by something else. |
| **Concentration** | The quantity of a chemical dissolved in a stated volume of solution. Measured in grams per litre. |
| **Accuracy** | How close a quantitative result is to the true or actual value. |
| **Outlier** | A measured result that is very different to other measured values or from what you would expect. This suggests the value is wrong. |
| **Mean value** | A type of average found by adding up all the measurements then dividing by the number of measurements. |
| **Best estimate** | The value in which you have the most confidence (usually the mean average) |
| **Range** | The difference between the highest and lowest in a set of measurements. |
| **Real difference** | The difference between the two mean values is real if their values do not overlap the ranges. |
| **Chemical change / reaction** | A change that forms a new chemical. |
| **Atom** | The smallest particle of an element. |
| **Hydrocarbon** | A compound of hydrogen and carbon only. Usually from crude oil. |
| **Oxidation** | A reaction that adds oxygen to a chemical. |
| **Combustion** | When a chemical reacts rapidy with oxygen releasing energy. |
| **Reactants** | The chemicals on the left hand side of the equation (ingredients). |
| **Products** | The chemicals on the right hand side of the equation (what is made). |
| **Chemical formulae** | A way of describing a chemical using symbols and numbers that gives us information about the type and number of atoms. |
| **Word equation** | A summary of a chemical reaction in words. |
| **Chemical equation** | A summary of the chemical reaction using symbols and numbers. |
| **Conservation of atoms** | All the atoms present at the beginning of a reaction will be there at the end. No new atoms are created or destroyed. |
| **Conservation of mass** | The total mass of chemicals is the same at the beginning and end of a reaction.  |
| **Correlation** | A relationship that affects an outcome. |
| **Cause** | A factor that affects an outcome. |
| **Factor** | A variable that changes and may affect something else. |
| **Outcome** | A variable that changes as a result of something else changing. |
| **Efficiency** | The percentage of energy supplied to a machine that is usefully transferred by it. |
| **Catalytic converter** | A device fitted to a vehicle exhaust that changes the waste gases into less harmful ones. |
| **Reduction** | A reaction that removes oxygen from a chemical. |
| **Regulations** | Rules that can be enforced by an authority. |
| **Wet scrubbing** | A process used to remove pollutants from flue gases. |