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| **Key Word** | **Definition** |
| **Element** | Simplest form of material. All elements are listed in the Periodic Table. Made of only one type of atom. |
| **Atom** | The smallest particle of an element. The atoms in an element are the same. Different elements have different atoms. |
| **Period** | A row in the periodic table. |
| **Group** | A column in the periodic table containing similar elements. |
| **Properties** | The physical or chemical characteristics of something (a chemical or a material) |
| **Reactivity** | How likely it is that a chemical reaction will take place. Highly reactive = strong tendency to react.  Reactivity series of metals – metals in order of reactivity from most reactive to least reactive (they should know this from KS3) |
| **Chemical Symbols** | The letters that represent elements in the periodic table. The first letter is always a CAPITAL. The subsequent letter(s) are always small letters. E.g. Mg = Magnesium |
| **Formula / Formulae** | A way of describing a chemical that uses symbols (letters) for atoms. Also gives numbers of different atoms.  e.g. H2O (two hydrogen atoms and one oxygen) |
| **Equation** | Can be a word equation or symbol equation. |
| **Displacement** | A more reactive chemical will kick out a less reactive chemical from a compound. A compound is made of more than one type of atom. |
| **Halogen** | Group 7 of the periodic table. Make salts when reacted with metals. Examples include Chlorine or Iodine. |
| **Alkali** | A compound that dissolves in water to give a pH higher than 7 (8-14). Examples include Sodium Hydroxide which is made when sodium reacts with water.  Turns universal indicator blue/purple coloured. |
| **Spectra** | Different colours that make up light. Flame spectra tell us what elements are found in a compound. |
| **Spectroscopy** | The use of instruments to study spectra. This can tell us about what chemicals are in different compounds. |
| **Electron** | A type of particle found in an atom. It is responsible for giving the atom a charge. Electrons orbit the nucleus of an atom and are responsible for properties of an element. Electrons are negatively charged. |
| **Proton** | Protons are tiny particles found in the nucleus of an atom. They are positively charged. |
| **Neutron** | Neutrons are tiny particles found in the nucleus of an atom. They have no charge. |
| **Electron Orbits / Shell** | The way that the electrons are arranged around the nucleus. |
| **Crystalline** | A material with molecules or ions that are lined up in a regular way as in a crystal. |
| **Molten** | Melted |
| **Electrolysis** | Splitting up a chemical compound into its elements using electricity. |
| **Ions** | An electrically charged atom or group of atoms. (This means that they have either lost or gained electrons.) |
| **Chemical Species** | The different chemical forms that an element can take. For example chlorine can be an atom, molecule or an ion. Each form has distinct properties. |
| **Metal** | Elements on the left side of the periodic table. Metals have characteristic properties. |
| **Compound** | A chemical made of more than one type of element (or atom), joined together by a reaction. |
| **Mixture** | More than one element or compound but not joined chemically together. |